**Periodic Trends Worksheet**

* + - 1. What is the atomic radius? Explain the trend.
			2. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.
			3. Rank the atoms by increasing atomic size and explain a. Br, Rb, Kr b. Se, Br, Cl c. Mg, Ca, Li
			4. Rank the atoms by increasing ionization energy and explain: a. K, Ca, Rb b. Kr, Br, Rb c. Sn, In, Sb
			5. Why does fluorine have a higher ionization energy than iodine?
			6. Rank the atoms by increasing electron affinity and explain: a. Sb, Bi, P b. Mg, Na, S
			7. What is the difference between electron affinity and ionization energy?
			8. Rank the atoms by increasing electronegativity and explain: a. N, F, Be b. S, Te, O
			9. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.
			10. Why do elements in the same family generally have similar properties?

**Periodic Trends Worksheet**

1. What is the atomic radius? Explain the trend.
2. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.
3. Rank the atoms by increasing atomic size and explain a. Br, Rb, Kr b. Se, Br, Cl c. Mg, Ca, Li
4. Rank the atoms by increasing ionization energy and explain: a. K, Ca, Rb b. Kr, Br, Rb c. Sn, In, Sb
5. Why does fluorine have a higher ionization energy than iodine?
6. Rank the atoms by increasing electron affinity and explain: a. Sb, Bi, P b. Mg, Na, S
7. What is the difference between electron affinity and ionization energy?
8. Rank the atoms by increasing electronegativity and explain: a. N, F, Be b. S, Te, O
9. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.
10. Why do elements in the same family generally have similar properties?

**Periodic Trends Worksheet**

* + - 1. What is the atomic radius? Explain the trend.
			2. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.
			3. Rank the atoms by increasing atomic size and explain a. Br, Rb, Kr b. Se, Br, Cl c. Mg, Ca, Li
			4. Rank the atoms by increasing ionization energy and explain: a. K, Ca, Rb b. Kr, Br, Rb c. Sn, In, Sb
			5. Why does fluorine have a higher ionization energy than iodine?
			6. Rank the atoms by increasing electron affinity and explain: a. Sb, Bi, P b. Mg, Na, S
			7. What is the difference between electron affinity and ionization energy?
			8. Rank the atoms by increasing electronegativity and explain: a. N, F, Be b. S, Te, O
			9. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.
			10. Why do elements in the same family generally have similar properties?

**Periodic Trends Worksheet**

* + - 1. What is the atomic radius? Explain the trend.
			2. Rank the following elements by increasing atomic radius: carbon, aluminum, oxygen, potassium.
			3. Rank the atoms by increasing atomic size and explain a. Br, Rb, Kr b. Se, Br, Cl c. Mg, Ca, Li
			4. Rank the atoms by increasing ionization energy and explain: a. K, Ca, Rb b. Kr, Br, Rb c. Sn, In, Sb
			5. Why does fluorine have a higher ionization energy than iodine?
			6. Rank the atoms by increasing electron affinity and explain: a. Sb, Bi, P b. Mg, Na, S
			7. What is the difference between electron affinity and ionization energy?
			8. Rank the atoms by increasing electronegativity and explain: a. N, F, Be b. S, Te, O
			9. Rank the following elements by increasing electronegativity: sulfur, oxygen, neon, aluminum.
			10. Why do elements in the same family generally have similar properties?