

$$2. \quad V = 100 \text{ mL} \div 1000 \\ = 0.1 \text{ L}$$

$$c = 0.865 \frac{\text{mol}}{\text{L}}$$



$$n_{\text{NaOH}} = c \times V \\ = 0.1 \cancel{\text{L}} \times 0.865 \frac{\text{mol}}{\cancel{\text{L}}} \\ = 0.0865 \text{ mol}$$

$$M_{\text{NaOH}} = 39.997 \frac{\text{g}}{\text{mol}}$$

$$m_{\text{NaOH}} = 0.0865 \text{ mol} \times \frac{39.997 \text{ g}}{\text{mol}} \\ = 3.46 \text{ g}$$

1. Mass out 3.46g of NaOH into a clean dry beaker.
2. Dissolve it in a minimum amount of deionized water.
3. Transfer (with rinsings) to a 100 mL volumetric flask.
4. Fill up to the line with deionized water.
5. Cap & shake.