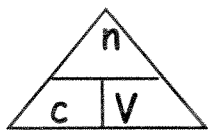


Making Solutions Warmup 2



$$C_c V_c = C_d V_d \quad \text{mass} \begin{array}{c} \xrightarrow{\div M} \\ \xleftarrow{\times M} \end{array} \text{mole}$$

1. How would you prepare 500.0 mL of a 1.75 mol/L HCl solution starting with an 8.61 mol/L stock solution?

Math

Lab Steps

2. How would you prepare 100.0 mL of a 0.865 mol/L solution of sodium hydroxide?
Hint: Sodium hydroxide is a solid

Math

Lab Steps