

- Draw Lewis structures for the following molecular compounds:
 - $F_2(g)$
 - $H_2O(l)$
 - $CH_4(g)$
 - $PCl_3(s)$
 - $H_2S(g)$
 - $SiO_2(s)$
- Draw Lewis structures for each of the following molecules:
 - $H_2(g)$
 - $O_3(g)$
 - $OF_2(g)$
 - $NF_3(g)$
 - $N_2H_2(g)$
 - $P_2H_4(g)$
- Draw a Lewis structure for the following polyatomic ion: $OH^-(aq)$
- Distinguish between bonding electrons and lone pairs.
- Are the following pairs of atoms more likely to form ionic compounds or covalent bonds?
 - sulfur and oxygen
 - iodine and iodine
 - calcium and chlorine
 - potassium and bromine
- Use a Lewis dot diagram to explain the formula for nitrogen, N_2 .
 - Draw the Lewis structure for nitrogen.
 - Nitrogen is a fairly inert (unreactive) gas. Explain this, referring to the bond involved.
- Illustrate the formation of each of the following molecular compounds, using Lewis dot diagrams and Lewis structures.
 - HCl
 - NH_3
 - H_2S
 - CO_2
- Is it correct for the structural diagram of H_2S to be written $H-H-S$? Explain using a diagram.

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