

1. State the four quantum numbers and the property they describe and the rules that they follow.
2. Define the following and give an example
   1. Pauli Exclusion Principal
   2. Aufbau Principal
   3. Hund’s Rule
3. Write out all the possible orbitals for n=4
4. Write out the diagonal aid to help you remember the order for filling up orbitals.
5. Write the full and shorthand electron configurations for the following: a. Mg b. Ti c. Fe2+
6. Write the condensed orbital box diagram for the following:
7. As3- b. Cu c. Cr
8. Review your models of the atom!!