**Concentration Practice Quiz /20** don’t forget sig fig, units, proper set-up, concluding statements

1. Two different clear, colourless liquids were gently heated in an evaporating dish. Liquid A left no residue, while liquid B left a white residue. Which liquid was a solution, and which was a pure substance. Explain. [2]
2. Explain, with a picture, how water is able to dissolve an ionic solid such as calcium chloride. [2]
3. Calculate the molar concentration of these aqueous solutions:
   1. 10.3 g table sugar (C12H22O11) in 100.0 mL solution [3]
   2. 50.0 mL of 0.150 mol/L NaOH diluted to 0.200 L with water [3]
4. How would you prepare these aqueous solutions (be specific)?
   1. 2.00 L of 0.457 mol/L KBr from solid KBr [3 calculation, 2 description]
   2. 500.0 mL of 0.0756 mol/L LiNO3 from 2.912 mol/L LiNO3 [3 calculation, 2 description]

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